

## CONTINGENCY SYLLABUS GENERAL CHEMISTRY B

**Instructor:** Willetta Greene Johnson, Ph. D.

**Contact:** ~~Office: Cudahy Science Hall Rom 322 773-508-3537~~

[wgreene@luc.edu](mailto:wgreene@luc.edu)

**Who am I:** A chemical physicist (Stanford University/ U. of Chicago) interested in surface interactions and mildly interested in (1) thermodynamics (2) swarm intelligence; (3) producer, composer, orchestrator, arranger and conductor. In Spring 2019, The Cincinnati Symphony performed a work of mine originally commissioned by The Chicago Sinfonietta; I also guest conducted in that concert. One of my songs was doubly tracked on a Grammy award winning CD in 2004. It has since been covered by seven other groups.

**Lectures: Up to and including Jan 31, 2022:** Lectures may be viewed **asynchronously** via Panopto. They are easily assessed through Lesson Pages on SAKAI.

**Zoom Friday Lecture ID:** **ID 816 5521 0911** (also for Office Hours)

**Zoom Office Hour: W (Jan 26) 9:20 A–10:10 A CST (contingency only)**

1. **Prerequisites:** CHEM 101 and 111, or 105 and completion of Math 118 or higher with a grade of C- or better. **Co-requisite:** 112.

### 2. Required Text:

- **Chemistry, the Central Science.** 14<sup>th</sup> ed. Theodore L. Brown, et. al. Boston: Pearson Prentice Hall: 2011 ISBN: **978-0134414232**. (*Mastering Chemistry* is NOT required).

**Chemistry 102 Course Packet**, authored by the instructor.

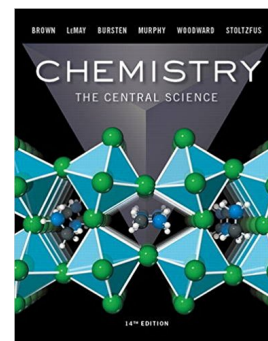
This lecture packet is available online at

<https://store.cognella.com/60064-4B-002>

- The course packet will be mailed to you within a few days of ordering, and you'll have immediate online access to 30% once order is completed.

**Discussions:** occur synchronously according to the following schedule:

Section	ZOOM ID	DAY and Time
<b>Sync Lect.</b>	<b>816 5521 0911</b>	<b>F: 9:20 A –10:10 A</b>
<b>010</b>	<b>884 8578 7653</b>	<b>F: 12:30 A – 1:25 P</b>
<b>011</b>	<b>875 2747 6813</b>	<b>F: 1:40 P – 2:30 P</b>
<b>012</b>	<b>848 5667 7289</b>	<b>F: 2:45 P – 3:35 P</b>



**Due to the large number of students / sections that are matriculated through this course yearly, there can be *absolutely no alteration* of this schedule.**

**Course Description:** A study of chemical principles and generalizations with emphasis on the development of a scientific attitude and an understanding of the fundamental concepts of chemistry.

**Calculators:** A scientific calculator is sufficient. Calculators cannot be shared while exams are in progress and their cases/covers must be removed. Be familiar with the calculator and the status of its batteries. The student is responsible for having a working calculator in lecture and for exams.

**Cell Phone<sup>1</sup> Policy:** It is forbidden to video/audio record lecture (except Loyola University staff). Stills of the board may be taken by permission. ***During exams, cell phone, wireless devices, and unauthorized materials are strictly forbidden; subject to device confiscation and dismissal from exam.***

<sup>1</sup> All technology, smart phone, tablets, laptops, surface pros, ... **Violations will be treated as instances of academic dishonesty** (see page 8)

**SAKAI Connection:** The syllabus, homework assignments for the semester, discussions, and discussion answers will be posted at the following website: [www.luc.edu](http://www.luc.edu), look under LINKS, click on **Sakai**). Students possessing a Loyola email address should be able to access this site.

### Objective of this course in grander detail:

By the conclusion of this course, the student should experience the following outcomes:

1. Understand the fundamental principles of physical chemistry
2. Acquire a knowledge base of basic terminology and classifications
3. Apply concepts creatively as well as solve multi-tiered problems
4. Know how to rank, estimate, analyze and critically evaluate a range of models
5. Gain a broader understanding of the role of chemistry in human endeavor
6. Appreciate the collaborative and global effort of the scientific enterprise

Specifically the engaged student should improve in her or his ability to

▪ **Comprehend the following concepts:**

- Intermolecular forces categories and contributing factors
- Kinetics, reaction rate, Arrhenius equation, reaction mechanism, rate limiting step
- Dynamic equilibrium and Equilibrium constants
- Titrations, Buffers, pH, Lewis and Arrhenius A/B models, Solubility product
- Complex ions and octahedral Crystal Field theory
- Electrochemical cells (batteries, fuel cells, transduction, respiration)
- Spontaneity vs. nonspontaneity, entropy, cyclic processes, and free energy

▪ **Identify reagents and general chemical processes:**


- Identify acids, bases, acidic and basic salts, buffers, solubility rules (Chemistry101)
- write appropriate net ionic prototypical rxns in aqueous solution
- be able to closely estimate and accurately calculate pH
- equilibrium constant K and how it predicts spontaneity status of a reaction
- complex ion, Lewis bases, geometry of “simpler” complex ions and compounds.




- **Assess outcome feasibility:** estimate energy and entropy of chemical processes
- **Work and exchange ideas with others:** cordially solve weekly group problems
- **Appreciate the impact of chemistry:** realize better how chemistry impacts life processes, technology, local, and global issues.
- **Contribute constructively:** as a science-literate, ethically responsible citizen and voter.

Later this semester, you will receive an emailed invitation to assess me via the **SmartEvals**. This survey provides a thorough diagnostic of how effectively students think the instructor realized the objectives boxed above, as well as the value of the course and other contextual experiences. This survey will be emailed to you from the Office of Institutional Effectiveness and will be available for approximately a one-week window only.

As student raters, please be aware that the results of your ratings for this class will be included as part of the information used to make decisions about promotion/tenure/salary increases for this instructor. Fairness to both the individual and the institution require *accurate and honest* answers.

## Chemistry 102 Schedule of Topics

Date	Day	Topic	Chapter	Pages (approx)
1/19 JAN	W	Classification of matter	11	444 - 452
1/21	F	Liquids & Interm. Forces	11	452-466
1/24	M	Boiling point Trends	12	472-500
1/26	W	Kinetics & Chemical Rate	14	574-582
1/28	F	Integrated Rate Law; <b>Example:</b> Radioactive Decay	14 21	582-590, 591-600, 913-916 (902-905)
1/31	M	Arrhenius Eq'n; k vs. T Rxn mechanisms	14	601-606 607-615
2/2, 4 FEB	W, F	Rate limiting step	14	615-622
2/7	M	Gas phase Equil'm	15	628-637
2/9	W	Const. K; Equil. Table	15	632-644; 644-650
2/11	F	van't Hoff equation	15	648-656
2/14	M	Le Chateliér's Principle	16	657-660
2/16	W	Dissoc'n $K_a, K_b$ ; Eq. Table	16	664-676; 678-680
2/18	F	A/B categories...SA, SB	16	670-679
2/21	M	WA, WB ; $K_a$ & $K_b$	16	681-695
2/23	W	REV 11,12,14-16	(~ 25 min)	OPTIONAL
2/25	F	Exam 1: Ch.11, 14-16	In person	
2/28	M	Conjugates; $pK_a + pK_b = 14$ A/B Salts, Binary acids	16	681-695 694-699, (702-704)
3/2 MAR	M	Buffers	17	724-729
3/4	W	Titrations SA/SB; SA/WB; SB/WA	17	730-739
3/6	F	Titration, continued	17	730-739
3/7 - 3/12		SPRING BREAK		
3/14	M	$K_{sp}$ and ppt equil'm Common Ion Effect	17	740-743 744-746
3/16 	W	Complex Ions (Lewis Acid/Base)	23 17	1002-1006 756-758
3/18	F	Ligands, Nomenclature	23	1007-1016

Date	Day	Topic	Chapter	Pages
3/21	M	Crystal Field Theory	23	1020-1026, 1030-3
3/23	W	Review: 16,17, 23 ( $K_{sp}$ )	~25 min	Optional, no handout
3/25	F	Exam 2: Ch.16,17, $K_{sp}$	In person	
3/28	M	Spont. And Temp. Statistical origin (physics)	19	812-818 821-828 (opt)
3/30	W	Entropy and 2 <sup>nd</sup> Law of Thermodynamics	19	818-821, 828-831
4/1	F	Gibb's Energy and Temp Gibbs Energy and K	19	831-838 838-841
4/4	M	Electrochemical Cell Spontaneity	20	868-874
4/6	W	$E_{red}^{\circ}$ , $E_{ox}^{\circ}$ , Spontaneity Work and $\Delta G$	20	868-874 876-879
4/8	F	Graphical Technique & spontaneity	20	Lecture/ Worksheets
4/11	M	Nernst Equation $E^{\circ}$ and Equil'm const.	20	880-886 886-892
4/13	W	Electrolysis / applications	20	893-896
4/14-4/18 APRIL	Th-M 	<b>EASTER BREAK</b> 		Nernst Equation $E^{\circ}$ and Equil'm const
4/20	W	Review: 19, 20, 23	~25 min	Optional, no handout
4/22	F	Exam 3: Ch. 19, 20, 23	In person	
4/25	M	Solutions; Solubility Units	13	524-532
4/27	W	Assorted unit conversions	13	533-540
4/29	F	Colligative Properties	13	541-558
4/30	SAT	<b>REVIEW FOR FINAL EXAM</b>	11-17, 19,20,23	<b>OPTIONAL</b>
5-4-2022	W	<b>COMMON FINAL EXAM: Ch. 11 - 17, 19, 20, 23</b>	7 PM - 9 PM CST Location tba	Format TBA Probably in person

## Representative Problems, End of Chapter Problems & Discussions:

Students who are making good progress should be able to solve, independently, most or all of the end-of-chapter problems in the text, listed below as assigned problems. A comprehensive review containing additional problems will be posted approximately one week before midterm exams; these reviews also serve for final exam.

**Discussions** (10% of grade) are held **ONLINE**. They should be attended. Assigned worksheet problems will be solved by groups of 3-4 students within the 50-minute discussion. Students must stay the entire period (unless otherwise instructed) and work on the problem(s) to earn up to 10% of grade. Solutions will be posted on SAKAI and graded discussions will be returned by the following discussion, or during the same week if an exam occurs in that week. The student is strongly encouraged to attempt all suggested problems (text-book and discussion) and contribute significantly to the group discussion activity. **The extent of group participation will be noted** and mentioned in any recommendation letters.

## Assigned Exercises:

chapter	page	Problems (*means more involved;** means unassigned reading may be required)
11	464	1, 2, 6, 7, 9, 15,17,19, 21, 23, 37, 39, 41, 47, 52, 54, 57, 61, 64, 85: Clausius Clapeyron Eq'n and 2 data points to find $\Delta H_{\text{vap}}$ .
12	514	11-13,15, 17, 32, 47 (look online), (guess: 4, 6, 7),
14	610:	3, 5, 7, 9, 12, 14, 19, 21 a - c, 23, 25 (sim. to 23), 27, 29, 33, 35, 37, 39, 41, 43, 46, 50 a-b (glucose is $\text{C}_6\text{H}_{12}\text{O}_6$ ) 57, 61, 63*, 67, 68, 73, 75, 91, 95, 99,105, *117.
21, only kinetics	938:	35 – 39a, *41, **43; <b>radioactivity</b> (radionuclide decay) is a 1 <sup>st</sup> order kinetic process $1 \text{ Becquerel (Bq)} = \text{s}^{-1}$ . $1 \text{ gray (Gy)} = 1 \text{ J per kg of tissue}$ . $1 \text{ rad} = 10^{-2} \text{ J/kg of tissue}$
15	656:	3, 5, 7, 9, 13, 15, 17, 23, 25, 26, 27, 28, 33a, 35, 37 ( $K_p = K_c$ why??), 39, 41, 45 find $[\text{Cl}_2]$ , next $PV = nRT$ , 49, 51, 53, 55, 57, Le Chateliér: 61, 63, 65, 68, *74a,c, 79, 81. <b>I don't overly emphasize <math>K_c</math> vs <math>K_p</math> but for practice, do #24, use formula 15.15 on page 631</b>
16	709:	1, 4, 5, 10, 15, 19, 21, 23, 25, 27, 29, 33, 37, 43, 45, 47, 53, 55, 57, 61, 65a-b, 69, 71, 73, 77, 79, 81, 83, 84, 85, 93, 109, 115, 118*; 95**, 99** **compares various A/B models % ionization = $x/x_0 \cdot 100\%$ ; $x_0$ = starting concentration, unit usually M
17	766:	1, 5, 6, 17 (% ionization $\leftrightarrow [\text{H}^+]$ ),19, 21, 23, 27, 29, 33, 35, 37, 41, 43, 47,10, 53, 55, 57, 61, 69, 73, 83, 12**, 67 a-b note: $K_{\text{sp}}$ , and $K_f$ needed in Prob. 67 concurrently.
19	839:	3, 4, 6, 9ab, 11, 13, 15, 21, 25, 35, 37, 41, 45, 47, 49, 53, 59, 61, 63, 65, 69, 73ab; 79, 83 ( $K_a$ ), 85, 88, 103, 96*: for each species, 1 <sup>st</sup> find $\Delta G^\ominus$ (appendix), 2 <sup>nd</sup> : then use $\Delta G = \Delta G^\ominus + RT \ln Q$ ; note: $\Delta G^\ominus = -RT \ln K_p$ . 75*, 92*, 106**
20	891:	4, 6, 7, 12, 13,15,17 part ii; 19, [23, 25 review balance redox] 29a-e, 31, 37, 39, 43, 53, 55, 57, 59, 61a, 63ab, 65, 67, 76ab, 84b, 88, 99, 102; 106* (organic-y)
21	936:	1, 5, 9, 11, 13,17, 21, 29, 35, 37, 47, 49, 55, 61 <b>extra</b> : 72, 74; 31**, 51*, 63*, 80*
23	1023:	15, 16, 23, 25, 27, 33-35, 37, 41, 43, 47, 55, 57, 59, 61, 63; <b>extra nerdy</b> : 73, 74, 78
13	559:	1, 3, 7, 11, 15, 17a, 22, 25, 27, 32, 37-39, 41, 43, 47, 57, 61, 67, 75, 77, 78

## Grading Scheme:

There are two grading schemes, and whichever one yields the higher grade will be employed after the final has been taken:

If all midterms went fairly well: 20% midterm, 10% discussion, 30% final

If one midterm not so good: 20% the other two, 10% discussion, 50% final

## Grading Scale:

<b>Grade Scale:</b>	<b>A</b> <b>≥ 91</b>	<b>A-</b> <b>88-90</b>
<b>B+</b> <b>85-87</b>	<b>B</b> <b>81-84</b>	<b>B-</b> <b>78-80</b>
<b>C+</b> <b>75-77</b>	<b>C</b> <b>71-74</b>	<b>C-</b> <b>67-70</b>
<b>D+</b> <b>64-66</b>	<b>D</b> <b>61-63</b>	<b>F</b> <b>&lt; 60</b>

## Examinations

Three hour exams and the final exam will be given on **February 25, March 25, April 22, and May 4**, respectively, also noted in the schedule. 90% of your course grade will be determined from these as explained further below. The other 10% will be determined from your discussion grades including excellent participation. The exams are cumulative, *i. e.*, may include material that has been queried on previous exams.

**All exams will be administered IN CLASS.**

## Final Exam:

The University sets the schedule for all final exams. The final will be held on:

**MAY 4, 2022 7:00 P-9:00 P CST**

The final exam is comprehensive and cumulative. You will have exactly 2 hours to complete the exam. Additional time will not be granted, even if you arrive late. There will be no make-up final exams given under any circumstance, and the exam will not be given early either. Instructors may not reschedule final exams for a class for another day and/or time during the final exam period. There can be no divergence from the posted schedule of dates for final exams.

Individual students who have four (4) final examinations scheduled for the same date may request to have one of those exams rescheduled. If a student reports having four final examinations scheduled for the same date, students should be directed to e-mail a petition to Lester Manzano, Assistant Dean for Student Academic Affairs, CAS Dean's Office ([lmanzan@luc.edu](mailto:lmanzan@luc.edu))

**Please note that the final examination *must* be taken. Failure to take the final exam will result in the grade "F".**



## Academic Honesty:

All students in this course are expected to have read and to abide by the demanding standard of personal honesty, drafted by the College of Arts & Sciences, which can be viewed at: <http://www.luc.edu/cas/advising/academicintegritystatement/>

A basic mission of a university is to search for and to communicate the truth as it is honestly perceived. A genuine learning community cannot exist unless this demanding standard is a fundamental tenet of the intellectual life of the community. Students of Loyola University Chicago are expected to know, to respect, and to practice this standard of personal honesty.

Academic dishonesty can take several forms, including, but not limited to cheating, plagiarism, copying another student's work, and submitting false documents.

Any instance of dishonesty (including those detailed on the website provided above or in this syllabus) will be reported to The Chair of The Department of Chemistry & Biochemistry who will decide what the next steps may be. As in the past, cheating will be SEVERELY dealt with—

- *minimally* costing the offender a grade of “zero” for the item that was submitted; this grade *cannot* be dropped. Additionally,
- incident is reported to both the Chemistry Department Chair and the Office of the CAS Dean.
- incident will be noted on one's transcript.
- depending upon the seriousness of incident, additional sanctions may be imposed.

## Loyola University Absence Policy for Students in Co-Curricular Activities (including ROTC):

Students missing classes while representing Loyola University Chicago in an official capacity (e.g. intercollegiate athletics, debate team, model government organization) shall be allowed by the faculty member of record to make up any assignments and to receive notes or other written information distributed in the missed classes.

Students should discuss with faculty the potential consequences of missing lectures and the ways in which they can be remedied. Students must provide their instructors with proper documentation (develop standard form on web) describing the reason for and date of the absence.

This documentation must be signed by an appropriate faculty or staff member, and it must be provided as far in advance of the absence as possible. It is the responsibility of the student to make up any assignments. If the student misses an examination, the instructor is required to give the student the opportunity to take the examination at another time. (<https://www.luc.edu/athletheadvising/attendance.shtml>)



### Accommodations for Religious Reasons

If you have observances of religious holidays that will cause you to miss class or otherwise effect your performance in the class you must alert the instructor **within 10 calendar days of the first class meeting of the semester** to request special accommodations, which will be handled on a case by case basis.

### Course Repeat Rule:

Effective Fall 2017 semester, students are allowed only THREE attempts to pass Chemistry courses with a C- or better grade. The three attempts include withdrawals (W). After the second attempt, the student must secure approval for a third attempt. Students must come to the Chemistry Department, fill out a permission to register form or print it from the Department: <http://www.luc.edu/chemistry/forms/> and obtain a signature from the Undergraduate Program Director, Asst. Chairperson, or Chairperson in Chemistry and Biochemistry. A copy of this form is then taken to your Academic Advisor in Sullivan to secure final permission for the attempt.

### Laboratory:

Chemistry 112, the general chemistry laboratory course, should be taken concurrently with the lecture course in general chemistry. The lecture and the laboratory courses are graded independently. Students should first consult the Chemistry Department Bulletin opposite the wall facing the chemistry office for information. Then, if they still have unresolved issues, they should contact **Dr. Katrina Banaku**, the administrator of the laboratories.

**Statement of Intent:** By remaining in this course, students are agreeing to accept this syllabus and to abide by the guidelines outlined in the document. Students will be informed should there be a necessary change to the syllabus.

**Intellectual Property:** All lectures, notes, Power Points and other instructional materials in this course are the intellectual property of the professor. As result, they may not be distributed or shared in any manner, either on paper or virtually without my written permission.

Lectures may not be recorded without my written consent; when consent is given, those recordings may be used for review only and may not be distributed. Also recognizing that your work is your intellectual property, I will not share or distribute your work in any form without your written permission.

**Class Conduct:** One important aspect of a Jesuit education is learning to respect the dignity, rights, and opinions of others. Please respect others by

- (1) allowing all classmates the right to voice their opinions without fear of ridicule
- (2) not using profanity or making objectionable (gendered, racial or ethnic) comments, especially comments directed at a classmate.

**Special Circumstances—Receiving Assistance:** Students are urged to contact me should they have questions concerning course materials and procedures. If you have any special circumstance that may have some impact on your course work, please let me know so we can establish a plan for assignment completion. If you require assignment accommodations, please contact me early in the semester so that arrangements can be made with Services for Students with Disabilities (SSWD) (<http://www.luc.edu/sswd/>).

### Student Support Resources:

- ITS HelpDesk 773-508-4487
  - [helpdesk@luc.edu](mailto:helpdesk@luc.edu)
- Library
  - Subject Specialists:  
<http://libraries.luc.edu/specialists>
- Services for Students with Disabilities
  - <http://www.luc.edu/sswd/>
- Writing Center
  - <http://www.luc.edu/writing/>
- Ethics Hotline- 855.603.6988
  - <http://luc.edu/sglc/aboutus/>

### SI - Faculty Partnership

There are **online** Supplemental Instruction (SI) study sessions available for this course. SI sessions are led by an SI leader, who is a student that has recently excelled in the course. Session attendance is open to all, and while it is voluntary, it is extremely beneficial for those who attend weekly. Times and locations for the SI session can be found here: [www.luc.edu/tutoring](http://www.luc.edu/tutoring). Students who attend these interactive sessions solve and discuss pertinent problems and concepts and share study / test-taking strategies. Research shows students whom regularly attend sessions have higher grades at the end-of-the-semester and more deeply understand course concepts than those who do not. Students are asked to arrive with their Loyola ID number, lecture notes, and textbook. This year, Spring 2022, our SI is **Sara Kennady** [skennady@luc.edu](mailto:skennady@luc.edu).

### SEATING POLICY / COVID

For the safety of everyone attending this course, there will be **assigned seating** and **attendance** will be taken. This will assist with contact tracing should it become necessary. We trust that students will understand and cooperate with this policy.

## Advanced Studies Recommendation Protocol

Later on, you may require a letter of recommendation (LOR) for graduate school, medical school, or the like. If I am chosen among your recommenders, the following policy ensues:

1. **Deadline for LOR (letter of recommendation) requests: Feb 1<sup>st</sup> of the application year.**
2. Student must generally possess GPA of 3.5 or above. However, a student might be considered if she or he presents a **written explanation** that reveals an exceptional circumstance accounting for a lower grade point average.
3. Student must provide attached in one email, a document listing his/her correct GPA, contact information, deadline(s), and also all chemistry, biology and physics courses and labs that the student has taken—in the following format (or Committee format, if you are applying through committee):
  - a. **GPA**
  - b. reliable, current email and telephone # that student checks *regularly*
  - c. **DEADLINE**
  - d. Table with header: course taken, instructor, grade

**Example:**

Course	Semester/year	Instructor	Grade
Chemistry 101	Fall / 2022	Dr. WGJ	B+
Biology 210	Spring / 2023	Dr. James LoDolce	A-

- e. If applying through Committee, *student initiates* process that culminates in a link being sent to me with which to submit LOR.
  - f. If applying “outside the Committee”—see items 5, 6 below, a list of all schools of the applicant and **ALL of their DEADLINES**.
  - g. Regarding part f: all cover forms, application packages, envelopes in one binder, folder, or otherwise secure containment, with like items paper-clipped together.
4. I'd love to read your personal statements, even in rough draft form. It tells me something about you and helps me to shape a recommendation. This article is not required, but I recommend it.
  5. **It is STRONGLY recommended that the student apply through the Loyola Pre-Health Advisory Committee.** Well-regarded by the medical /dental/ pharmaceutical community, the Committee's voice of endorsement will increase the merit of the student's application. Their method also assures that the student's personal statement is strong and well written. If the student applies via Committee, I will be contacted.
  6. **APPLICATIONS OUTSIDE COMMITTEE:** If a student who I can recommend elects to apply apart from the Pre-Health Advisory committee, she/he must perform steps 2-4 and email me at [wgreene@luc.edu](mailto:wgreene@luc.edu). Online LOR uploading protocols (AMCAS, PTCAS, Interfolio, etc.) are **STRONGLY** preferred.

Due to the volume of requests, your **LOR** won't be started until all items in 3 are fulfilled